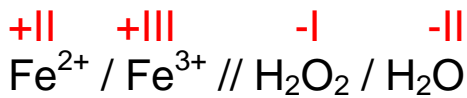
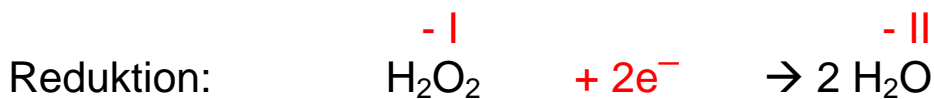
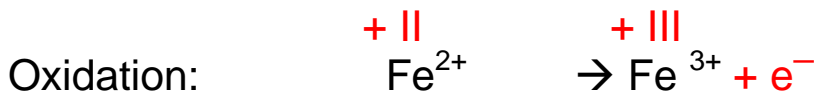


Aufgabe 1:  $\text{FeSO}_4 + \text{H}_2\text{O}_2 \rightarrow \text{Fe(III)-Ionen} + \text{H}_2\text{O}$

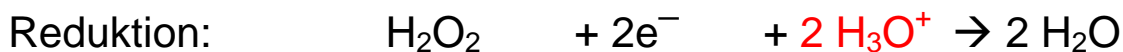
### Schritt 1: Bestimmung der Redoxpaare:



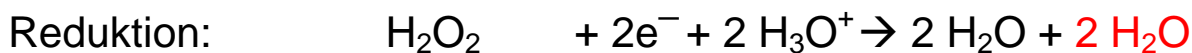
### 2: Aufstellen der Teilgleichungen incl. Elektronen



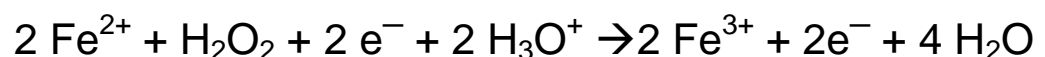
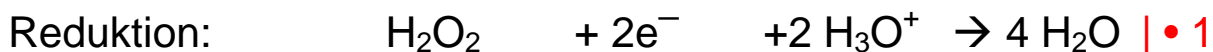
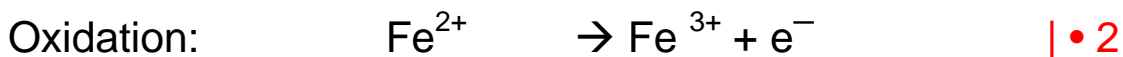
### Schritt 3: Ladungsausgleich



### Schritt 4: Stoffausgleich



### Schritt 5: kgV bilden und Gesamtgleichung aufstellen



Kürzen ergibt:  $2 \text{Fe}^{2+} + \text{H}_2\text{O}_2 + 2 \text{H}_3\text{O}^+ \rightarrow 2 \text{Fe}^{3+} + 4 \text{H}_2\text{O}$